Chemistry Formula Sheet

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CHAPTER 9

Redox Reactions

Oxidation number

$$\text{oxidation}=\text{loss of electrons (ON increases)}; \qquad \text{reduction}=\text{gain (ON decreases)}$$

RULES

$$\text{F}=-1;\ \text{O}=-2\ (\text{peroxide}\ -1,\ \text{OF}_2\ +2);\ \text{H}=+1\ (\text{metal hydride}\ -1);\ \sum\text{ON}=\text{charge}$$

Balancing and titrations

N-FACTOR

$$n\text{-factor}=\text{change in ON per formula unit}; \qquad \text{KMnO}_4:\ 5\ (\text{acid}),\ 3\ (\text{neutral}),\ 1\ (\text{strong base});$$

$$\text{K}_2\text{Cr}_2\text{O}_7:\ 6$$

EQUIVALENCE

$$N_1V_1 = N_2V_2; \qquad \text{meq of oxidant}=\text{meq of reductant}$$

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