Chemistry Formula Sheet

Chapter by chapter, topic by topic. Written by the PadhoNa team for quick revision.

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CHAPTER 18

Electrochemistry

Cells and Nernst equation

$$E^\circ_{\text{cell}} = E^\circ_{\text{cathode}} - E^\circ_{\text{anode}}$$

NERNST

$$E = E^\circ - \frac{0.0591}{n}\log Q \quad (298\ \text{K})$$

$$\Delta G^\circ = -nFE^\circ_{\text{cell}}, \qquad E^\circ_{\text{cell}} = \frac{0.0591}{n}\log K$$

Conductance

$$\kappa = \frac{1}{\rho} = G\,\frac{l}{A}, \qquad \Lambda_m = \frac{\kappa \times 1000}{M}$$

KOHLRAUSCH

$$\Lambda^\circ_m = \nu_+\lambda^\circ_+ + \nu_-\lambda^\circ_-; \qquad \alpha = \frac{\Lambda_m}{\Lambda^\circ_m}$$

STRONG ELECTROLYTES

$$\Lambda_m = \Lambda^\circ_m - A\sqrt{C}$$

Electrolysis

FARADAY'S LAWS

$$w = ZIt = \frac{E}{F}It, \qquad \frac{w_1}{w_2} = \frac{E_1}{E_2}$$

$$1\ \text{F} = 96500\ \text{C mol}^{-1}$$

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